

Chemistry - Peterson  
Date:

Name:  
Period:

**Lecture 10B: Molar Mass**

**What is gram atomic mass?**

One mole of an \_\_\_\_\_ = molar mass of the element

The unit for molar mass is \_\_\_\_\_.

What is the mass of 1 mol of each of these elements?

- 1 mol of C = \_\_\_\_\_,
- 1 mol of Sr = \_\_\_\_\_,
- 1 mol of Cu = \_\_\_\_\_,

**What is gram formula mass?**

One mole of an ionic \_\_\_\_\_ = molar mass of that compound

Example 1: 1 mole of  $K_2CO_3$  = \_\_\_\_\_ g of  $K_2CO_3$

K:

C:

O:

The subscripts tell you the number of \_\_\_\_\_!

Example 2: 1 mole of  $Ca_3(PO_3)_2$  = \_\_\_\_\_ g of  $Ca_3(PO_3)_2$

Ca:

P:

O:

**Summary:**

Now let's try a dimensional analysis problem...

Example #3: How many moles are in  $4.2 \times 10^{24}$  grams of  $C_2H_6$ ?

Summary: